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Tuning Photophysics and Nonlinear Absorption of Bipyridyl Platinum(II) Bisstilbenylacetylide Complexes by Auxiliary Substituents

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ABSTRACT: The photophysics of six bipyridyl platinum(II) bisstilbenylacetylide complexes with different auxiliary substituents are reported. These photophysical properties have been investigated in detail by UV−vis, photoluminescence (both at room temperature and at 77 K) and transient absorption (nanosecond and femtosecond) spectroscopies, as well as by linear response time-dependent density functional theory (TD-DFT) calculations. The photophysics of the complexes are found to be dominated by the singlet and triplet π,π* transitions localized at the stilbenylacetylide ligands with strong admixture of the metal-to-ligand (MLCT) and ligand-to-ligand (LLCT) charge-transfer characters. The interplay between the π,π* and MLCT/LLCT states depends on the electron-withdrawing or -donating properties of the substituents on the stilbenylacetylide ligands. All complexes exhibit remarkable reverse saturable absorption (RSA) at 532 nm for nanosecond laser pulses, with the complex that contains the NPh2 substituent giving the strongest RSA and the complex with NO2 substituent showing the weakest RSA.

INTRODUCTION

Photophysical and photochemical properties of platinum complexes have been of intense interest for researchers for several decades. These interests arise from the unique structural and spectroscopic properties of the platinum complexes and their promising potential applications in the fields of organic light-emitting diodes (OLEDs), low power photon upconversion sensitizers, dye-sensitized solar cells (DSSCs), DNA intercalators, photodynamic therapy (PDT), and so forth. Among the reported complexes, platinum diimine bis(acetylide) complexes are of particular interest. Systematic investigations of the photophysics and potential applications of these complexes have been reported by several groups. In 1994, Che and co-workers reported their study on the Pt(phen)(C=C=CC6H5)2 complexes and proposed that the emissive state for these complexes has a metal-to-ligand charge-transfer character (MLCT). Following that, comprehensive work by Eisenberg’s group discovered that the lowest occupied molecular orbital (LUMO) of the Pt(II) diimine complexes could be adjusted by variation of the substituents on diimine, while the highest occupied molecular orbital (HOMO) was slightly affected by the variation of the aryacetylide ligands. In recent years, Castellano and co-workers demonstrated the interplay of the MLCT and π,π* states in Pt(II) diimine complexes bearing more conjugated aromatic rings, that is, 1-pyrene, 1-anthracene, and 1-phenylacetylene on the acetylide ligands. In accordance with these results, our group recently revealed the admixture of the MLCT and π,π* characters in the lowest excited states of a 2,2′-bipyridine Pt(II) complex bearing 2-(benzothiazol-2-yl)-9,9-diethyl-7-ethynylfluorene ligands.

In addition to the interesting photophysical studies, the potential applications of the Pt(II) diimine complexes in OLEDs, molecular photochemical devices, and as vapor-luminescent materials and nonlinear absorbing materials have been explored. A detailed nonlinear absorption study reported by our group on the 2,2′-bipyridine Pt(II) complex bearing 2-(benzothiazol-2-yl)-9,9-diethyl-7-ethynylfluorene ligands revealed that this complex exhibited extremely large ratios of the excited-state absorption relative to that of the ground state in the visible spectral region and large two-photon absorption (TPA) in the near-IR region. This discovery is quite exciting; however, the two-photon absorption cross sections of this complex are still moderate compared to those of the best organic two-photon absorbing materials. In order to improve...
the TPA of the Pt(II) bipyridyl complexes in the near-IR region, we are interested in introducing the stilbene component to the acetylide ligand because the TPA of stilbene and its derivatives has been extensively investigated.\textsuperscript{16} The ultimate goal is to explore whether the TPA of the Pt(II) bipyridyl complexes could be improved by incorporation of the stilbene component while retaining the strong excited-state absorption of the Pt(II) bipyridyl complexes.

Schanze’s group has reported the photochemistry and photophysics of four platinum complexes bearing stilbenylacetylide ligands.\textsuperscript{17} It is found that the triplet excited states of these complexes are localized on one of the 4-ethynylstilbene ligands, and the coordination geometry can affect the extent of delocalization through the metal center. However, the effects of the substituents on the 4-ethynylstilbene ligands on the ground- and excited-state properties of the platinum diimine bistilbenylacetylide complexes have not been explored, and the nonlinear absorption characteristics of these complexes have not been investigated either.

In this study, we synthesized a series of platinum diimine bistilbenylacetylide complexes with different substituents on the stilbene motif (Scheme 1, complexes 1–6) and systematically investigated their photophysics and excited-state absorption. The nonlinear transmission performances of these complexes were evaluated at 532 nm using nanosecond laser pulses. Although the ultimate goal of this project is to enhance the TPA of the Pt(II) complexes in the near-IR region, the linear optical properties of these complexes and their dependence on the substituents are in close relationship to the nonlinear absorption of the complexes. As such, we first focus on understanding how the
auxiliary substituents on the stilbenylacetylide ligand influence the ground-state and excited-state characteristics of the complexes. To reach this goal, both photophysical experiments and time-dependent density functional theory (TD-DFT) calculations were carried out. To the best of our knowledge, it is the first time that such an effect has been explored for Pt(II) complexes bearing substituted stilbenylacetylide ligands. It is also the first time that the excited-state absorption and nonlinear transmission of the Pt(II) bis(stilbenylacetylide) complexes is systematically investigated.

## EXPERIMENTAL SECTION

Synthesis and Characterization. All of the reagents and solvents for synthesis were purchased from Aldrich Chemical Co. or Alfa Aesar and used as received, unless otherwise stated. Silica gel for chromatography was purchased from Sorbent Technology (60 Å, 230–400 mesh, 500–600 m²/g, pH: 6.5–7.5). Complexes 1–6 were characterized by ¹H NMR, electrospray ionization mass spectrometry (ESI-MS), and elemental analyses. Ligands 3-L–6-L were characterized by ¹H NMR and elemental analyses. Each intermediate was characterized by ¹H NMR. ¹H NMR was obtained on a Varian Oxford-VNMRSpectrometers (300, 400, or 500 MHz). ESI-MS analyses were performed at a Bruker BioTOF III mass spectrometer. Elemental analyses were conducted by NuMega Resonance Laboratories, Inc. in San Diego, California.

4,4′-Di(S,9-diethyltridecan-7-yl)-2,2′-bipyridine (21) was synthesized following the literature procedure. ¹H NMR was synthesized by the reaction of K₂PtCl₄ with 21 in refluxing aqueous HCl solution. Compounds 7, 9, 11, 13, and 15 were synthesized by the Heck reaction. Compound 18 was synthesized by the Wittig reaction. Compound 19 was synthesized by the Ullmann reaction from 18. The Sonogashira coupling reaction of 7, 9, 11, 13, 15, and 19 with ethynyltrimethylsilane or 2-methyl-3-buten-2-ol followed by hydrolysis with K₂CO₃ or KOH in i-PrOH afforded ligands 1-L–6-L. The synthesis of 1-L–6-L was all previously reported, except for 3-L, whose synthesis is described in the Supporting Information. The ¹H NMR data for compounds 21 and 22 are also provided in the Supporting Information. The target complexes 1–6 were obtained by reaction of compound 22 with the respective stilbenylacetylide ligand using disopropylamine as the base and CuI as the catalyst. The synthetic procedure and characterization data for 1–6 are provided in the Supporting Information.

DFT Calculations. The ground- and excited-state properties of complexes 1–6 were studied using DFT and linear response TD-DFT. All calculations (ground-state geometry optimization, electronic structure, excited states, and optical spectra) were performed using the Gaussian 09 quantum chemistry software package. The geometries of all molecules were optimized for the ground state (closed-shell singlet S₀). For calculations of emission spectra, the geometry was optimized in the lowest-energy excited singlet (S₁) and triplet (T₁) states. All procedures were done utilizing the hybrid long-range corrected CAM-B3LYP functional. The LANL08 basis set was used for the heavier Pt atom, while the remaining atoms were modeled with the 6-31G* basis set. The chosen method represents one of the currently most accurate DFT functionals and basis sets that have already shown good agreement with experimental data for different organometallic complexes. All calculations have been performed in solvent using the CPCM reaction field model as implemented in Gaussian 09. Dichloromethane (CH₂Cl₂, εᵣ = 8.93) was chosen as a solvent for consistency with the experimental studies. As was found for several organometallic complexes, inclusion of the solvent into calculations is very important to reproduce experimental optical spectra.

For computations of emission energies, the excited-state geometry optimization has been performed with analytic TD-DFT gradients using the long-range corrected functional CAM-B3LYP. For the absorption spectra of the complexes, the 40 lowest singlet optical transitions were considered to reach the transition energies of ∼4.9 eV. Each spectral line obtained from the TD-DFT calculation was broadened by a Gaussian function with the line width of 0.1 eV to match the experimentally observed homogeneous broadening. The fluorescence energies were determined by calculating vertical transition energies for the optimized lowest singlet excited state (S₁) geometries (both excited-state geometry optimization and vertical transitions have been performed with the TD-DFT). The phosphorescence energies were calculated by first optimizing the lowest triplet-state geometry followed by the vertical triplet excitations calculated via TD-DFT.

In order to analyze the nature of the singlet and triplet excited states, natural transition orbital (NTO) analysis was performed. NTOs are obtained via separate unitary transformations applied to the occupied (hole) and the unoccupied (electron) orbitals. Such unitary transformation diagonalizes the transition density matrix obtained from TD-DFT calculations and gives the best representation of the electron excitation in single-particle terms. In other words, this method offers the most compact representation of the transition density as a single pair of an electron–hole excitation from the ground state. Here, we refer to the unoccupied NTO (to which an electron is excited) as the “electron” transition orbital and the occupied NTO (from which an electron was transferred) as the “hole” transition orbital. Note that electron and hole NTOs are not the same as unoccupied and occupied molecular orbitals (MOs) in their ground state. Upon photoexcitation, the excitonic couplings (Coulomb interaction between the photoexcited electron–hole pair) mix the ground state MOs, so that the representation of an excitation via the pair of ground-state MOs is not valid. In contrast, electron and hole NTOs obtained from TD-DFT calculations allow for representation of the excited-state electronic density. NTOs shown in this paper were produced with the isovalue of 0.02 and visualized with the GaussView 5.1 graphical software.

Photophysical Measurements. The solvents used for photophysical experiments were spectroscopic grade, which were purchased from VWR International and used as received without further purification. An Agilent 8453 spectrophotometer was used to record the UV–vis absorption spectra in different solvents. A SPEX fluorolog-3 fluorometer/phosphorimeter was used to record the steady-state emission spectra in different solvents. The emission quantum yields were determined by the relative actinometry method in degassed solutions, in which a degassed 1 N sulfuric acid solution of quinine bisulfate (Φₑm = 0.546, λₑm = 347.5 nm) was used as the reference. The laser system for the femtosecond transient absorption measurement was described previously. Briefly, the 800 nm laser pulses were produced at a 1 kHz repetition rate (fwhm = 110 fs) by a mode-locked Ti:sapphire laser (Hurricane, Spectra-Physics). The output from the Hurricane was split into pump (85%) and probe (10%) beams. The pump beam (800 nm) was sent into an optical paramagnetic amplifier (OPA-400, Spectra Physics) to obtain 310, 330, and 350 nm excitation sources.
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Electronic Absorption. The UV–vis spectra of complexes 1–6 are shown in Figure 1. The absorption of all complexes obeys Lambert–Beer’s law in the concentration range of our study (1 × 10⁻⁶ to 1 × 10⁻⁴ mol/L), suggesting that no ground-state aggregation occurs in this concentration range. The major absorption bands of complexes 1–6 (Figure 1b) resemble those of their ligands (shown in Figure 1a), which are exemplified in Figure 1a and b for 1, 4, 5, and 6, indicating that these bands arise from the stilbenylacetylide ligands. This assignment is supported by the molar extinction coefficients of these bands, which are at least double those of their corresponding ligands. The red-shift of these bands suggests delocalization induced by the dₓ orbitals of platinum. The absorption of complexes 3–5 in the region of 300 and 375 nm features well-resolved vibronic structures, which is indicative of the σπ,π* transitions localized on the stilbenylacetylide ligands. However, at the wavelengths above 380 nm, a broad shoulder that is absent in the ligands’ absorption spectra is observed. Compared to the major absorption bands centered at ~340–350 nm, this shoulder exhibits a significant negative solvatochromic effect (as illustrated in Figure 2a), which is consistent with the other Pt(II) diimine complexes reported in the literature₁⁻⁹ and implies a charge-transfer nature of this absorption band. In contrast, the absorption spectra of 1, 2, and 6 are broad and red-shifted relative to the major absorption bands of 3–5. The vibronic structures are not clearly observed for 1, 2, and 6. Absorption of complexes 1 and 6 were also found to exhibit a negative solvatochromic effect, as illustrated in Figure 2c for 1. All of these features lead to the attribution of the absorption of 1, 2, and 6 partially to charge-transfer processes, probably mixing ligand-to-ligand charge transfer

\
\[ \varepsilon_T = \frac{\Delta OD_T}{\Delta OD_S} \]

where \( \Delta OD_S \) is the minimum of the bleaching band and \( \Delta OD_T \) is the maximum of the absorption band in the TA spectrum and \( \varepsilon_S \) is the ground-state molar extinction coefficient at the wavelength of the bleaching band minimum. After the \( \varepsilon_T \) value was obtained, the \( \Phi_T \) could be obtained by the relative actinometry, in which SiNc in benzene was used as the reference (\( \varepsilon_{100} = 70000 \text{ M}^{-1} \text{ cm}^{-1} ; \Phi_{100} = 0.20 \)).

Nonlinear Transmission. The reverse saturable absorption of complexes 1–6 was characterized by a nonlinear transmission experiment at 532 nm using a Quantel Brilliant laser. The pulse width of the laser was 4.1 ns, and the repetition rate was set at 10 Hz. The complexes were dissolved in CH₂Cl₂. The concentration of the sample solutions was adjusted to obtain a linear transmission of 80% at 532 nm in a 2 mm thick cuvette. The experimental setup and details are similar to those reported previously. A 40 cm plano-convex lens was used to focus the beam to the center of the 2 mm thick sample cuvette.

Results and Discussion

Molecular Geometries. The optimized geometries of complexes 1–6 are reported in the Supporting Information (Table 1s). Optimization of the ground singlet state of the complexes 1–6 with DFT resulted in their essentially planar geometry. The stilbenyl components of the stilbenylacetylide ligands were twisted relative to the Pt coordination plane by ~1°, while substituent groups stayed in the stilbenyl plane (except for the NPh₂ group, which was twisted in the propeller-like shape). The average Pt–N bond length was 2.087 Å and was not affected by the stilbenylacetylide substituent group. Pt–C bond lengths were ~1.940 Å and were very slightly affected by the substituent groups (in the fourth decimal digit).

\( (E < 1 \text{ μJ/pulse}) \) and to a second harmonic generator (Super Tripler, CSK) to obtain a 400 nm beam (\( E \approx 2 \text{ μJ/pulse} \)). The probe beam was focused into a rotating CaF₂ crystal or a sapphire crystal for white light continuum generation between 350 and 750 nm or between 400 and 800 nm. The flow cell (Starna Cell Inc. 45-Q-2, 0.9 mL volume with a 2 mm path length), pumped by a Fluid Metering RhSY Lab pump (Scientific Support Inc.), was used to prevent photodegradation of the sample. After passing through the cell, the continuum was coupled into an optical fiber and input into a CCD spectrograph (Ocean Optics, S2000). The data acquisition was achieved using in-house LabVIEW (National Instruments) software routines. The group velocity dispersion of the probing pulse was determined using nonresonant optical Kerr effect (OKE) measurements. Sample solutions were prepared at a concentration needed to have absorbance of \( A \approx 0.6–0.8 \) at the excitation wavelength.

The nanosecond transient absorption (TA) spectra and decays were measured in degassed solutions on an Edinburgh LP920 laser flash photolysis spectrometer. The third harmonic output (355 nm) of a Nd:YAG laser (Quantel Brilliant, pulse width = 4.1 ns, repetition rate set at 1 Hz) was used as the excitation source. Each sample was purged with argon for 30 min before each measurement. The triplet excited-state absorption coefficient \( (\varepsilon_T) \) at the TA band maximum was determined by the singlet depletion method. The following equation was used to calculate the \( \varepsilon_T \).

\( \varepsilon_T = \frac{\Delta OD_T}{\Delta OD_S} \)

where \( \Delta OD_S \) is the minimum of the bleaching band and \( \Delta OD_T \) is the maximum of the absorption band in the TA spectrum and \( \varepsilon_S \) is the ground-state molar extinction coefficient at the wavelength of the bleaching band minimum. After the \( \varepsilon_T \) value was obtained, the \( \Phi_T \) could be obtained by the relative actinometry, in which SiNc in benzene was used as the reference (\( \varepsilon_{100} = 70000 \text{ M}^{-1} \text{ cm}^{-1} ; \Phi_{100} = 0.20 \)).
(LLCT) and MLCT from $d\pi$ (platinum) to $\pi^*$ (diimine). However, considering the large molar extinction coefficients of 1, 2, and 6, the lowest-energy absorption band in these three complexes should have dominant contribution from the acetylide ligand $\pi,\pi^*$ transition.

In general, the major absorption bands of all substituted stilbenylacetylide complexes are red-shifted relative to that of complex 4, and a similar pattern is observed in the ligands' absorption. This result agrees with the previous study of the push–pull conjugated oligomers and suggests that electronic structures of the Pt complexes discussed in this work can be effectively tuned by varying the substituents at the 4′-position of the stilbenylacetylide ligands, which consequently influences the photophysical properties.

These trends, as well as the overall spectral assignment are supported by the linear response TD-DFT calculations. Calculated absorption spectra (Figure 1c) show a substantial agreement with the experimental data, as can be seen in Figures 1 and 2. Although theoretically predicted absorption energies (Figure 1c) are blue-shifted compared to experimental ones, they clearly reproduce the experimental trends of absorption spectra and are found to be highly dependent on solvent polarity (Figure 2b). Experimental and theoretical absorption properties are summarized and compared in Table 1. Unlike 1, 2, and 6, complexes 3–5, which are substituted with weak electron–donor/acceptor groups, exhibit a lower-intensity shoulder in the spectral region above 380 nm. As resolved by the DFT studies, the shoulder involves two excitations with mixed MLCT and LLCT character, while the higher-energy peak has dominant ligand $\pi,\pi^*$ character.

The excited-state character of the studied complexes can be easily seen from the NTO plots, which represent the optical transition as a single pair of the excited electron and the empty hole orbitals as a result of an electron photoexcitation from the ground state. NTOs are shown in Table 2 and in the Supporting Information. Thus, the lowest-energy transitions of complexes 3–5 involve a transfer of the Pt d electrons from the top of the valence band (see the diagram in Figure 3) and the stilbenylacetylide ligand to the bipyridine (bpy) moiety. The MLCT/LLCT characters of these transitions explain their relatively small oscillator strengths (intensities of peaks in absorption spectra). Due to unequal interaction of the ligand field with the participating Pt d orbitals of different symmetry, these excited states are noticeably separated (by $\sim 0.5$ eV) from higher-energy states, which explains the appearance of the well-separated, low-intensity shoulder in the absorption spectra of these compounds. The second absorption band of compounds 3–5 can mainly be described as the $\pi,\pi^*$ intraligand transition within the stilbenylacetylide ligands with an admixture of the

<table>
<thead>
<tr>
<th>Complex</th>
<th>$\lambda_{obs}/\text{nm}$</th>
<th>$\epsilon_{max}/\text{M}^{-1}\text{cm}^{-1}$</th>
<th>$\lambda_{max}/\text{nm}$ (ex. state)</th>
<th>$f_{osc}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>401 (69330), 298 (48900)</td>
<td>389 (S1; 1.6529), 375 (S2; 1.8782), 353 (S3; 0.3870)</td>
<td>389 (S1; 1.8782), 353 (S3; 0.3870)</td>
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<tr>
<td>2</td>
<td>380 (90250), 292 (40080)</td>
<td>383 (S1; 1.4325), 367 (S2; 1.4423), 347 (S3; 1.1505)</td>
<td>383 (S1; 1.4325), 367 (S2; 1.4423), 347 (S3; 1.1505)</td>
<td></td>
</tr>
<tr>
<td>3</td>
<td>420 (14580), 363 (78730), 349 (88280), 333 (74500)</td>
<td>378 (S1; 0.8110), 364 (S2; 0.9787), 333 (S3; 1.7741)</td>
<td>381 (S1; 0.8476), 363 (S2; 0.8989), 329 (S3; 1.7507)</td>
<td></td>
</tr>
<tr>
<td>4</td>
<td>410 (16480), 360 (87300), 344 (97850), 325 (77800)</td>
<td>375 (S1; 0.5847), 363 (S2; 0.9974), 333 (S3; 1.1800)</td>
<td>381 (S1; 0.7001), 363 (S2; 0.9190), 333 (S3; 1.5416)</td>
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<tr>
<td>5</td>
<td>410 (14980), 360 (89250), 349 (103840), 335 (87230)</td>
<td>375 (S1; 0.5847), 363 (S2; 0.9974), 333 (S3; 1.1800)</td>
<td>381 (S1; 0.7001), 363 (S2; 0.9190), 333 (S3; 1.5416)</td>
<td></td>
</tr>
<tr>
<td>6</td>
<td>385 (100480), 305 (62130)</td>
<td>389 (S1; 1.3019), 374 (S2; 1.4423), 353 (S3; 0.3870)</td>
<td>384 (S1; 0.7001), 363 (S2; 0.9190), 333 (S3; 1.5416)</td>
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</tr>
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</table>

$\lambda_{obs}$ is the absorption wavelength, and $\epsilon_{max}$ is molar extinction coefficient. $\lambda_{max}$ is the calculated wavelength corresponding to the transition between the ground and excited states of interest (the number of the excited state is shown in parentheses), and $f_{osc}$ is the calculated oscillator strength for the corresponding excitations.
MLCT/LLCT character (see NTOs for the third excited state, S3, in Table 2). This band is more sensitive to the type of substituent and red-shifted in case of the stronger electron donor or acceptor attached to the stilbenyl moiety.

In contrast, the lowest-energy absorption bands of complexes 1, 2, and 6 are predicted to have strong contribution of the \( \pi,\pi^* \) intraligand transition, with some mixture of the MLCT and LLCT characters (see Table 2). Significant ligand \( \pi,\pi^* \) character of these transitions gives rise to the intensity of the lowest-energy absorption band of these complexes, as can be seen in Figure 1c. NTOs in Table 2 for complexes 1 and 6 illustrate the delocalized nature of the lowest-energy excited

<table>
<thead>
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<th>Excited state number and properties</th>
<th>Hole</th>
<th>Electron</th>
</tr>
</thead>
<tbody>
<tr>
<td>S1 ( f_{osc} = 1.6529 ) 3.19 eV 389 nm</td>
<td><img src="image1" alt="Image" /></td>
<td><img src="image2" alt="Image" /></td>
</tr>
<tr>
<td>S1 ( f_{osc} = 0.6476 ) 3.26 eV 381 nm</td>
<td><img src="image3" alt="Image" /></td>
<td><img src="image4" alt="Image" /></td>
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<tr>
<td>S1 ( f_{osc} = 1.7507 ) 3.77 eV 329 nm</td>
<td><img src="image5" alt="Image" /></td>
<td><img src="image6" alt="Image" /></td>
</tr>
<tr>
<td>S1 ( f_{osc} = 1.3019 ) 3.19 eV 389 nm</td>
<td><img src="image7" alt="Image" /></td>
<td><img src="image8" alt="Image" /></td>
</tr>
<tr>
<td>S3 ( f_{osc} = 1.5416 ) 3.53 eV 352 nm</td>
<td><img src="image9" alt="Image" /></td>
<td><img src="image10" alt="Image" /></td>
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</table>

Table 2. NTOs\(^a\) Representing Transitions That Correspond to the First and the Second Lower-Energy “Bands” for Complexes 1, 4, and 6\(^b\)

\(^a\)Note that excited-state NTOs differ from the ground-state MOs and rather can be considered as the linear combination of the ground-state MOs that contribute to a given excited state. The NTOs were calculated based on the optimized ground-state geometry. Each electronic transition \( S_n \) is presented as a single electron–hole pair excitation from the occupied “hole” NTO to the unoccupied “electron” NTO. \(^{b}\)The excited-state number, corresponding oscillator strengths, and excitation energies are shown for each optical transition.
states that contribute to the absorption band, which results in larger oscillator strengths for these transitions. The higher-lying excitations can rather be characterized as the LLCT transitions from the stilbenylacetylide to the bpy moiety mixed with MLCT character, which explains their lower intensity.

In general, the shape of the absorption spectra of the discussed Pt complexes with the stilbenylacetylide ligands depends on the interplay between the states with charge-transfer character (LLCT and MLCT) and the intraligand \( \pi, \pi^* \) transitions. Thus, the complexes with the weak donor/acceptor substituents (3–5) exhibit a MLCT/LLCT shoulder in the lower-energy region of the spectra. In the cases when the stilbenylacetylides are substituted with a stronger electron-donor/acceptor substituent (1, 2, and 6), the lower-energy MLCT/LLCT states strongly mix with the intraligand \( \pi, \pi^* \) transitions, which causes larger transition dipole moments and stronger absorption intensity. The ground-state MO diagram shown in Figure 3 provides a better understanding of the effect of the substituent on energy levels and absorption spectra of the Pt complexes discussed herein. The highest occupied orbitals are delocalized over the stilbenylacetylide ligands (L–\( \pi \)) and have noticeable Pt-d contribution, while the lowest unoccupied orbitals are predominantly delocalized over the bipyridine ligand (bpy–\( \pi^* \)), except for complex 1. Addition of the electron-donating groups, such as OCH\(_3\) (5) and NPh\(_2\) (6), increases the ground-state dipole moment and raises the energies of the HOMOs. For complex 6, HOMO and HOMO–1 can be characterized as the pure ligand \( \pi \) (L–\( \pi \)) orbitals with very little Pt-d admixture. Addition of electron-withdrawing substituents, such as NO\(_2\) (1) and CHO (2), lowers the energies of the unoccupied levels, which lie very close to or even below (as in complex 1) the bpy–\( \pi^* \) level. Therefore, addition of stronger electron-donor or stronger electron-acceptor groups leads to a smaller HOMO–LUMO gap (red shift in absorption spectra) and change in the character of the lowest-energy absorption bands.

Addition of polar solvents compensates for the static and the transition dipole moments induced by the strong electron-donating/accepting groups and stabilizes the donor/acceptor electronic levels. This is equivalent to shifting HOMOs to lower energies, while the LUMO localized on the stilbenyl

and the emission data are summarized in Table 3. The emission band(s) of 2–5 at room temperature are mirror images to those of their respective major absorption band(s) with Stokes shifts smaller than 80 nm, while the emission of 1 and 6 at room temperature is broad and structureless, with Stokes shifts of 129 and 107 nm, respectively. The emission lifetimes of these complexes could not be detected by our spectrometer due to the very short lifetime (<5 ns). These features suggest that the observed emission for complexes 1–6 at room temperature emanates from the singlet excited state of the stilbenylacetylide ligand. At 77 K, the emission of 1–6 appears at 625–667 nm (Figure 4c), which is assigned to the phosphorescence of the complexes. However, due to the very weak signal, the phosphorescence lifetimes have not been determined.

In previous studies\(^7\) of compound 4, the singlet \( ^1\pi, \pi^* \) state has been proposed to relax predominantly through the intersystem crossing into the triplet \( ^3\pi, \pi^* \) state, which explains the observed relatively weak fluorescence. The triplet \( ^3\)MLCT emission has also been observed in a similar series of complexes but shown to undergo a quick nonradiative decay through the

**Figure 3.** Ground-state molecular orbital energy diagram for complexes 1–6 calculated in dichloromethane using the CAM-B3LYP functional and the LANL08 basis set for the Pt atom and the 6-31G* basis set for other atoms.

**Figure 4.** Normalized emission spectra of (a) ligands 1-L, 4-L, 5-L, and 6-L and (b) complexes 1–6 in dichloromethane at room temperature; (c) complexes 1–6 in butyronitrile glassy matrix at 77 K (phosphorescence).
relaxation of the stilbene ligand(s) into the cis conformation.\textsuperscript{17} The energy levels of the $^3\pi,\pi^*$ states relative to the $^1$MLCT states have been shown to strongly depend on the type of substituent on the 4,4′-position of the bpy ligand because of a strong influence of the substituent on the $^3$MLCT energies.\textsuperscript{41} Similar to previous studies, we found a strong effect on the relative positioning between the $^3\pi,\pi^*$ and $^1$MLCT states that is caused by the strength of the electron-donor/acceptor substituent on the stilbene.

To elucidate the nature of emission of complexes 1–6, their spectra are compared to the emission of the corresponding ligands 1-L, 4-L, 5-L, and 6-L (Figure 4a). The emission of 4-L can be assigned to the $^1\pi,\pi^*$ transition by comparing to the trans-stilbene emission reported in the literature.\textsuperscript{42} The $^1\pi,\pi^*$ character of the singlet emission in this ligand is also well presented in the NTOs, as shown in Figure 5. Compound 5-L emits at a similar level as 4-L (therefore, its NTOs are not shown in Figure 5), while the 1-L and 6-L spectra are significantly red-shifted due to the presence of strong electron-accepting/donating groups. NTOs for the relaxed singlet excited-state geometries, corresponding to the fluorescence, show that strong electron-donating substituent NPh, delocalizes the hole NTO over both the stilbene and the substituent group, while showing more localized character of the electron NTO on the stilbene with insignificant electron density on the substituent group (see the bottom panel in Figure 5). The strong electron-withdrawing NO$_2$ group has an opposite effect, with the more delocalized character of the electron NTO than the hole. Such a difference in delocalization between the electron and hole NTOs explains the red shift in the fluorescence of ligands 1-L and 6-L compared to that of other ligands.

The emission spectra of complexes 1 and 6 very closely represent emission of their ligands, with a slight red shift, which is indicative of the predominant intraligand character of the photoluminescence in this range (see Figure 4). However, the width of the spectra significantly increases as a result of some interactions with Pt d orbitals, as well as strong electron–vibrational couplings. The other complexes emit at substantially lower energy compared to their corresponding ligands, which suggests a strong delocalization induced by platinum dπ orbitals, and significant contribution from $^1$MLCT/$^1$LLCT, as discussed in the following paragraph.

Overall, TD-DFT calculations of the singlet and triplet emission reproduce the trends observed experimentally, as summarized in Table 3. NTOs for the relaxed singlet and triplet excited-state geometries (corresponding to fluorescence and phosphorescence, respectively) are illustrated in Figure 5 and compared with the NTOs contributing to the singlet emission of the corresponding stilbenylacetylide ligands. The singlet emission of complex 1 stems from the intraligand $^1\pi,\pi^*$ transition in one of the single stilbenylacetylide ligands, which becomes slightly twisted relative to the Pt coordination plane and, thus, hinders the $^1$MLCT character of fluorescence in this compound. Interestingly, the character of the NTOs contributing to singlet emission of this compound has a very different nature compared to that of NTOs contributing to the lowest-energy absorption (compare NTOs in Figure 5 with S1 in Table 2 for 1). Such a significant difference in NTOs is a sign of a strong excitation–vibration coupling, which is evident by a broadening of the fluorescence peak of complex 1 compared to that of 1-L shown in Figure 4. In all other complexes, NTOs contributing to fluorescence are strongly mixed with the $^1$MLCT from Pt to the bpy moiety, as well as LLCT between the stilbene and bpy. Strong charge-transfer character of the NTOs involved in the singlet emission explains the experimentally observed red shift in the emission maxima of these complexes versus the fluorescence energies of their ligands.

Table 3. Photoluminescence and Excited-State Absorption Parameters for Complexes 1–6 and Ligands 1-L, 4-L, 5-L, and 6-L

<table>
<thead>
<tr>
<th>Complex</th>
<th>$\lambda_{\text{em}}$/nm</th>
<th>$\Delta\tau$/ps</th>
<th>$\Phi_{\text{em}}$/nM</th>
<th>$\epsilon_{\text{em}}$/M$^{-1}$cm$^{-1}$</th>
<th>$\lambda_{\text{T,T}}$/nm</th>
<th>$\Phi_{\text{T}}$/$\text{nm}$</th>
<th>$\text{Thres}_{\text{emp}}$/nm</th>
<th>$\text{Thres}_{\text{emp}}$/nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>1-L</td>
<td>530 (0.012)</td>
<td>667</td>
<td>480 (385; 67825; 750 (401; 98760; 1.07)</td>
<td>495 (59)</td>
<td>1.70</td>
<td>497 (99)</td>
<td>663</td>
<td></td>
</tr>
<tr>
<td>2</td>
<td>454 (0.008)</td>
<td>659</td>
<td>510 (225; 184760; 0.079)</td>
<td>532 (77)</td>
<td>1.70</td>
<td>472 (7)</td>
<td>638</td>
<td></td>
</tr>
<tr>
<td>3</td>
<td>409 (0.017)</td>
<td>630</td>
<td>435 (48; 73; –)</td>
<td>465 (14)</td>
<td>1.70</td>
<td>449 (39)</td>
<td>638</td>
<td></td>
</tr>
<tr>
<td>4</td>
<td>425 (–)</td>
<td>625</td>
<td>460 (64; 73; –)</td>
<td>460 (69)</td>
<td>1.70</td>
<td>452 (69)</td>
<td>638</td>
<td></td>
</tr>
<tr>
<td>5</td>
<td>428 (0.009)</td>
<td>630</td>
<td>460 (73; –)</td>
<td>468 (205)</td>
<td>1.70</td>
<td>460 (205)</td>
<td>638</td>
<td></td>
</tr>
<tr>
<td>6</td>
<td>492 (0.003)</td>
<td>656</td>
<td>520 (198; 235510; 0.075)</td>
<td>526 (18)</td>
<td>1.70</td>
<td>463 (68)</td>
<td>668</td>
<td></td>
</tr>
</tbody>
</table>

In CH$_2$Cl$_2$. In BuCN glassy matrix. \(^3\pi,\pi^* \) transition to the bpy ligand because of a strong influence of the substituent on the $^3$MLCT energies.\textsuperscript{41}
excited state. Overall, variation of the substituent at the 4'-position of the stilbenylacetylide ligands has a significant influence on the emission energy and electronic state of the complexes.

**Transient Absorption.** The nanosecond and femtosecond transient absorption (TA) measurements for complexes 1−6 and ligands 1-L, 4-L, 5-L, and 6-L were carried out. From the TA experiment, we can obtain the spectral features and lifetimes of the singlet and triplet excited states. The spectral region where the excited-state absorption is stronger than the ground-state absorption can be identified from the positive absorption band(s) of the TA spectrum, while the excited-state lifetime can be deduced from the decay of the TA and the triplet excited-state quantum yield is determined by the relative actinometry.

The nanosecond TA spectra of complexes 1−6 in a CH$_3$CN/CH$_2$Cl$_2$ (v/v = 10/1) solution at zero delay after excitation are shown in Figure 6. All of the complexes exhibit strong TA signals that are well time-resolved, as illustrated by the time-resolved spectra of 1−6 in Supporting Information Figures S2–S7. The TAs of the complexes are significantly enhanced compared to those of their corresponding ligands (Figure S1 in the Supporting Information), indicating enhanced intersystem crossing (ISC) induced by the heavy-atom effect of Pt.

**Figure 5.** NTO plots for the singlet ($S_1$) and triplet ($T_1$) emissions (corresponding to the fluorescence and phosphorescence, respectively) for complexes 1, 4, and 6 and their corresponding ligands 1-L, 4-L, and 6-L. NTOs were obtained based on the optimized excited-state geometry to simulate the photoluminescence.

**Figure 6.** Nanosecond transient difference absorption spectra of complexes 1−6 in CH$_3$CN/CH$_2$Cl$_2$ (v/v = 10/1) immediately after 4.1 ns laser excitation. $\lambda_{ex} = 355$ nm, and $A_{355} = 0.4$ in a 1 cm cuvette.
Complexes 1 and 2 that bear strong electron-withdrawing substituents show broad absorption bands, while the shapes of the TA spectra of the other complexes resemble those of their corresponding ligands. For all complexes, a bleaching band was observed at λ < 450 nm, which is consistent with the position of the major absorption band in their UV–vis absorption spectra. The triplet lifetime deduced from the decay of the TA and the molar extinction coefficient of the triplet excited-state absorption determined from the singlet depletion method for 1, 2, and 6 are listed in Table 2. Similar to those trends observed from the UV–vis absorption and emission spectra, the bleaching band and TA absorption band are red-shifted for complexes 1, 2, and 6 in comparison to those of 3–5. The triplet excited-state lifetimes of 1, 2, and 6 are also significantly shorter than those of 3–5, which indicates different natures of the excited state that gives rise to the TA. For complex 4, the feature of the TA spectrum and the lifetime are quite similar to those reported by Schanze and co-workers for a similar Pt(II) bipyridyl bis(ethynylstilbene) complex, in which the transient is believed to be a stilbene-localized π,π* state with the double bond in the trans geometry. The short lifetime should be attributed to the rapid nonradiative decay induced by the C=C bond rotation in the stilbene moiety. The similar TA patterns for complexes 3 and 5 to those of their respective ligands and to 4 imply that the excited state that gives rise to the TA is the stilbene-localized π,π* state as well.

For complexes 1 and 2, the TA spectra are much broader, and a new absorption band appears at 750 nm for 1. With reference to the TA study reported by Castellano and co-workers for Pt terpyridyl acetylide complexes bearing π-conjugated arylacetylides, the near-IR TA band of 1 could possibly originate from the acetylide cation generated from the LLCT/ILCT (intraligand charge-transfer from the stilbene to the nitro group) processes. Therefore, the excited state that gives rise to the observed TA for 1 is tentatively assigned to the stilbene-localized π,π* state, mixed with LLCT/ILCT characters. We believe that the TA for 2 has a similar origin.

The femtosecond time-resolved TA spectra of 1–6 in a mixed CH3CN/CH2Cl2 solution are shown in Figure 7 and in the Supporting Information. The TA spectra of 2–5 resemble those of their nanosecond TA spectra and thus are assigned to the same transients as those contributing to the nanosecond TA, which could be from the singlet excited state or the triplet excited state in case the intersystem crossing is very fast.
absorption cross sections and deduce the ratios of the excited-state absorption cross sections relative to those of the ground state. This work is currently in progress.

**CONCLUSION**

A series of platinum bipyridyl platinum(II) bisstilbenylacetylide complexes with different auxiliary substituents on the stilbenylacetylide ligands were synthesized, and the photophysics of these complexes were systematically characterized by spectroscopic measurements and TD-DFT theoretical calculations. The absorption and emission characteristics of these complexes can be substantially adjusted by the auxiliary ligands. While the substitution of H on the 4′-position of stilbene by Br and OMe groups does not alter the photophysical properties of the complexes eminently, the absorption and the emission characteristics are significantly tuned by the CHO, NO₂, and NPh substituents. The lowest-energy absorption bands of 3–5 are dominated by the MLCT/LLCT transitions with an admixture of the π,π* character, while the lowest absorption bands for complexes 1, 2, and 6 are dominated by the ligand-centered π,π* transition with the admixture of MLCT/LLCT characters (in strong dependence of the solvent polarity). Emission shows very similar trends. The nanosecond and femtosecond transient difference absorptions of the complexes feature strong absorption in the visible spectral region, which can be attributed to a stilbene-localized π,π* state with the double bond in the trans geometry. The TA spectra of complexes 1 and 2 that bear strong electron-withdrawing substituents are very broad and extend to the near-IR region, suggesting contribution from the acetylide ligand cation that results from the LLCT/ILCT transitions. All complexes exhibit strong reverse saturable absorption for the nanosecond laser pulse at 532 nm, with complex 6 showing the strongest RSA due to the very strong excited-state absorption relative to that of the ground state at 532 nm. This makes complex 6 a very promising candidate for devices that require strong RSA.

**ASSOCIATED CONTENT**

*Supporting Information*

The synthetic procedures and characterization data for 11, 12, 3-L, and 1–6, characterization data for 21 and 22, optimized molecular structures for 1–6 via DFT calculations, ground-state molecular orbitals for 1, 4, and 6, and NTO plots for the lowest excited states of 1–6 calculated in CH₂Cl₂, NTO plots for the lowest excited states 1 and 4 calculated in toluene, nanosecond TA spectra of ligands 1-L and 4-L–6-L at zero time delay, nanosecond time-resolved TA spectra of 1–6, femtosecond time-resolved TA spectra of 2–6, 1-L, and 2–6–L, and the full ref 24. This material is available free of charge via the Internet at http://pubs.acs.org.

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**Notes**

The authors declare no competing financial interest.

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